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Worksheet
Answers

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Chemistry This chemistry video tutorial focuses on molarity and dilution problems. It shows you how to convert between molarity, grams, ...

Molarity Problems Worksheet Answer Key

Dilution Problems
Worksheet 1. How do you prepare a 250.-ml

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of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO_3 is used to make a 2.5 L dilution, what is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is

Molarity Problems Worksheet - Mrs Getson's Blog

Dilutions Worksheet -
Solutions 1) If I add 25

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mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$ (0.15 M)(125 mL) = x (150 mL) x = 0.125 M

2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$

Dilutions Worksheet

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Teacher Resources

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better of ... #27 This week, Hank elaborates on why Fugu can kill you by illustrating the ideas of solutions and discussing molarity, molality, and ...

Dilution Problems, Chemistry, Molarity ...

Solutions Worksheet Answers

The first three problems are questions regarding molarity and the others involve the dilutions formula

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above. Determine the number of grams of NaHCO_3 that are in one liter of a 2.1 M solution Determine the number of grams of NaNO_3 that are present in 500 mL of a 1.0 M solution.

Making Dilutions Worksheet - Liberty Union High School ...

Molarity and Serial
Dilutions Teacher
Handout Module
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most important skills in science is to calculate dilutions. In the lab, correct dilutions of solutions are critical to performing accurate science. In medicine, these skills are necessary to ensuring proper medicine dosage and treatment.

Molarity and Serial Dilutions Teacher Handout

Molarity and Dilutions .
9. Ion Concentration .

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10. Molarity Unit

Review # 1 . 11.

Molarity Unit Review #

2 . 12. Chemistry 11

Calculations Practice

Test # 1 . 13.

Chemistry 11

Calculations Practice

Test # 2 . Molarity

Worksheet # 1 . 1.

15.8 g of KCl is
dissolved in 225 mL of
water. Calculate the
molarity.

Molarity Worksheet

1

Page 14/26

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Calculate the molarity of 0.289 moles of Iron (III) Chloride, FeCl_3 , dissolved in 120 of 1000 FL What is the

molarity of 0.5 grams of sodium chloride, NaCl , dissolved to make 50 mL of solution? $M \times V = 1 .65$

Calculate the molarity of 734 grams of lithium sulfate, Li_2SO_4 , dissolved in 2,500 mL of solution. $Z 500$

Molarity WS - HN

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KEY

Solutions Review
Worksheet Answers 3
Partial Pressures &
Vapor Pressure: Crash
Course Chemistry #15
This week we continue
to spend quality time
with gases, more
deeply investigating
some principles
regarding pressure -
including Molarity
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answers.

Dilution Worksheets - Kiddy Math

Dilutions Worksheet -
Solutions 1) If 45 mL of
water are added to 250
mL of a 0.75 M K_2SO_4
solution, what will
the molarity of the
diluted solution be?

$$(0.75 \text{ M})(250 \text{ mL}) = M$$

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$2 (295 \text{ mL}) M_2 = (0.75 \text{ M})(250 \text{ mL}) = 0.64 \text{ M}$
(295 mL) 2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what

Dilutions Worksheet

W 329 - Everett

Community College

8) I have 345 mL of a 1.5 M NaCl solution. If I boil the water until the volume of the solution is 250 mL, what will the molarity of the solution

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be? Dilutions

Worksheet. 1) If I have 340 mL of a 1.5 M NaBr solution, what will the concentration be if I add 650 mL more water to it?

Dilutions Worksheet - Socorro Independent School District

Dilutions. Showing top 8 worksheets in the category - Dilutions. Some of the worksheets displayed

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is the total number of
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calculate the molarity
of a solution, divide the

Lab Math Solutions, Dilutions, Concentrations and Molarity

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Dilution Problems Worksheets - Lesson Worksheets

If I took 180 mL of that solution and diluted it to 500 mL, determine the molarity of the resulting solution.

Solution: 1) Calculate moles of NaF: $125.6 \text{ g} / 41.9 \text{ g/mol} = 3.00 \text{ mol}$.

2) Calculate moles in 180 mL of resulting

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solution: 3.00 mol in 1000 mL so $3 \times (180/1000) = 0.54$ mol in 180 mL. 3) Calculate molarity of diluted solution:

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