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## **2010 Ap Chemistry Solutions**

AP <sup>®</sup> CHEMISTRY 2010 SCORING GUIDELINES Question 2 (10 points) A student performs an experiment to determine the molar enthalpy of solution

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of urea,  $\text{H}_2\text{NCONH}_2$ . The student places 91.95 g of water at  $25^\circ\text{C}$  into a coffee cup calorimeter and immerses a thermometer in the water.

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## Questions

A solution of 0.100 M HCl and a solution of 0.100 M NaOH are prepared. A 40.0 mL sample of one of the solutions is added to a beaker and then titrated with the other solution. A pH electrode is used to obtain the data that are plotted in the titration curve shown above. (a) Identify the solution that was initially

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added to the beaker.

## **AP CHEMISTRY 2010 SCORING GUIDELINES (Form B)**

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## **AP Chemistry 2010 Free-Response Questions Form B**

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GUIDELINES (Form B) Question 1 (10 points) The compound butane,  $C_4H_{10}$ , occurs in two isomeric forms, n-butane and isobutane (2-methyl propane). Both compounds exist as gases at  $25^\circ C$  and 1.0 atm. (a) Draw the structural formula of each of the isomers (include all atoms). Clearly label each structure.

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AP<sup>®</sup> CHEMISTRY 2010 SCORING  
GUIDELINES Question 1 (10 points)  
Several reactions are carried out using

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AgBr, a cream-colored silver salt for which the value of the solubility product constant,  $K_{sp}$ , is  $5.0 \times 10^{-13}$  at 298 K. (a) Write the expression for the solubility product constant,  $K_{sp}$ , of AgBr.  $K_{sp} = [Ag^+][Br^-]$

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Exam Content. The AP Chemistry Exam covers a full-year introductory college course in chemistry with laboratory. The multiple-choice questions in Section I cover the breadth of the curriculum. Section II, the free-response part of the exam, includes three quantitative problems, one question on writing chemical reactions and predicting

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products,...

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AP<sup>®</sup> CHEMISTRY 2010 SCORING  
GUIDELINES (Form B) Question 5 (9  
points) A solution of 0.100 M HCl and a  
solution of 0.100 M NaOH are prepared.  
A 40.0 mL sample of one of the solutions



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is added to a beaker and then titrated with the other solution. A pH electrode is used to obtain the data that are plotted in the titration curve shown above.

### **AP® Chemistry 2010 Scoring Guidelines Form B - MAFIADOC.COM**

Properties of Solutions 2 <sup>3</sup>/<sub>4</sub>

miscible—When two or more liquids mix

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(ex. Water and food coloring)  $\frac{3}{4}$   
immiscible—When two or more liquids  
DON'T mix.—they usually layer if  
allowed to set for a while. (ex. Water and  
oil) Concentration Units  $\frac{3}{4}$  Molarity (M) =  
# of moles of solute per liter of solution;  
IS temperature dependent.

## **AP\* Chemistry PROPERTIES OF**

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## **SOLUTIONS**

AP\* Chemistry: 2008 Released Multiple Choice Exam NO CALCULATORS MAY BE USED Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified. Throughout the test the following symbols have the definitions

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specified unless otherwise noted. Part A

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AP Chemistry: Solutions - Chapter Summary and Learning Objectives. A

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solution - one substance dissolved in another - is a cornerstone of chemistry theory and experimentation. In this chapter, you'll learn the different types of solutions and what factors affect them.

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QUESTIONS . GO ON TO THE NEXT PAGE.

-6-(c) Calculate the concentration of urea, in mol/L, in the saturated solution at 20.°C. (d) The student also determines that the concentration of urea in a saturated solution at 25°C is 19.8 . M. Based on this information, is the dissolution of urea endothermic or ...

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