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### 12 Stoichiometry Chapter Test A

1:2. Which of the following is true about the total number of reactants and the total number of products and their action shown below  $C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$ . 9 moles of reactants chemically change into 11 moles of product.

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Amount of given substance A in moles multiplied by #1: Mole ratio from balanced equation multiplied by #2 Molar Mass from Periodic table equals amount of unknown substance B in grams.  $A \text{ (in moles)} \times \#1 \times \#2 = B \text{ (in grams)}$

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$A \text{ (in grams)} \times \#2 \times \#1 = B \text{ (in moles)}$  Mass to Mass Conversions. Amount of given substance A in grams multiplied by #2 Molar mass from periodic table multiplied by #1 Mole ratio from balanced equation multiplied by #3 Molar mass from periodic table equals amount of unknown substance B in grams.

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### Chapter 12 - Stoichiometry - Preston Treend

380 Chapter 12 Assessment continued CHAPTER 12 49. a. 2.36 g  $H_3PO_4$  b. 1.89 g  $CO_2$  50. a.  $5.70 \times 10^{21}$  atoms Zn b. 95.2 g Zn 51. a.  $7.0 \times 10^2$  L  $N_2$  b. no reagent in excess 52. a. 96.4% b. 45.0 g 53. 10.7 kg  $CaSO_4$  54. 224 L gas 55. a. Initially, the amount of NaCl formed increases as the amount of Na used increases. For this part of the curve, sodium is the limiting reagent.

### CHAPTER 12 Study Guide

ANS: The coefficients from the balanced equation are used to write mole ratios. The mole ratios relate the moles of reactants to the moles of product. By multiplying the number of moles of the reactant by the mole ratio, you can determine the number of moles of the product. DIF: L2 REF: p. 359 OBJ: 12.2.1 43.

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### Chapter 12 Stoichiometry Packet Answer Key

d. stoichiometry e. actual yield f. excess reagent Column A the substance that determines the amount of product that can be formed in a reaction the amount of product that forms when a reaction is carried out in the laboratory the calculation of quantities in chemical equations the ratio of the actual yield to the theoretical yield expressed as ...

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Chapter 12 - Stoichiometry - Standardized Test Prep - Page 417: 6. Answer. Work Step by Step. Statement I: Every stoichiometry calculation uses a balanced equation True. The chemical equation needs to be balanced before we can do any calculation Statement II: Because every chemical reaction obeys the law of conservation of mass True.

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